Ex. Problems:

Sodium metal reacts with oxygen to produce solid sodium oxide. How many grams of sodium must react to produce ________ g of sodium oxide?

When ________ g of hydrogen reacts with oxygen, how many grams of water are produced?

Actual Yield: amount of __________ produced when the __________ is performed in a __________

Theoretical Yield: amount of __________ expected to be __________ based on the ____________ and the amount of __________

Percent Yield

(__________ yield / __________ yield) x 100%
Percent Yield in Lab

actual yield of CO$_2$ = ____________  (from lab data)

Calculate theoretical yield from balanced equation:

\[
\text{NaHCO}_3 + \text{HCl} \rightarrow \text{NaCl} + \text{CO}_2 + \text{H}_2\text{O}
\]

?g CO$_2$ = 0.23 g NaHCO$_3$

% yield =

What is the % yield of carbon dioxide when __________ grams of propane are burned and __________ grams of carbon dioxide are collected?

The Chemistry Quiz

CR1.  CR2.  1.  2.

3.  4.  5.

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